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thermodynamics has much wider applicability. In material science, one is normally not that interested in heat and work, but interested more the state of matter and how things might change when mixed, heated, pressurized, etc. Some important effects are chemical reactions (such as oxidation), formation of solutions, phase transformations.

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From partial molar results (eq. (8.15)), the free energy of the solution is $G_{\text{soln}} = n_A G_{A0} + n_B G_{B0} + RT [n_A \ln X_A + n_B \ln X_B] + H_{\text{mix}} + n_A L_A \ln X_A + n_B L_B \ln X_B + RT [n_A \ln X_A + n_B \ln X_B]$. Because G_{A0} and G_{B0} are for constants that do not depend on subsequent increase in temperature, we need to solve for an

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increase in G_{soln} using Solve @ $H_1 \hat{=} 2L RT H n_A \text{Log}@X_A D + n_B \text{Log}@X_B D L == RT H n_A \text{Log}@X_A D + n_B \text{Log}@X_B D + H n_A + n_B L \text{Log}@PDL , PD 99P \emptyset E -\text{Log}@X_A D n_A ...$

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David R. Gaskell , David E. Laughlin CRC Press , Feb 7, 2003 -
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There's no excuse. Despite the fact that you have to read slowly and carefully to make sense of anything, I found this book practically useless as a tool for learning thermodynamics. Gaskell goes on and on and on and on and on about mundane details (the book could easily be 200 pages lighter), while glossing over the important points.

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